



A2 Physical Chemistry

Brønsted – Lowry theory of acids and bases

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Brønsted – Lowry theory of acids and bases

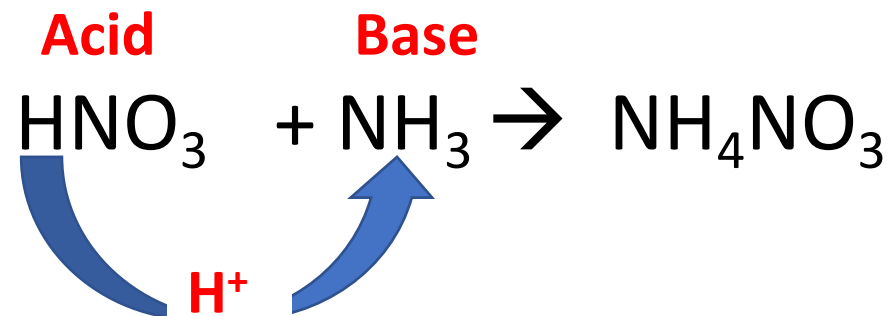
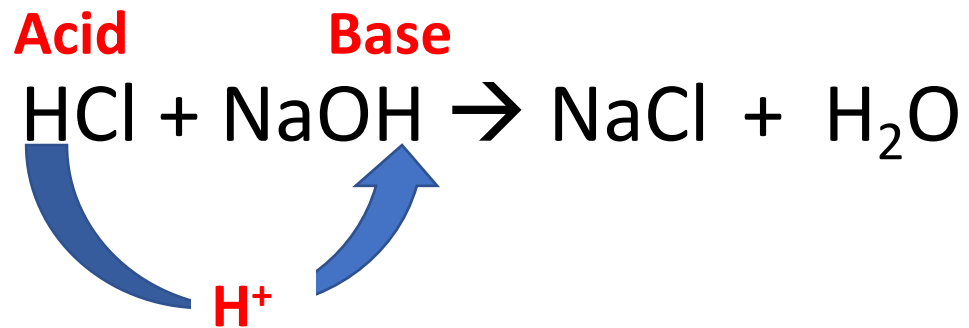
An acid is a proton (H^+) donor and a base is a proton acceptor.

Brønsted-Lowry acid-base reaction = a reaction involving the transfer of a proton

Bronsted – Lowry theory of acids and bases

An acid is a proton (H^+) donor and a base is a proton acceptor.

Bronsted-Lowry acid-base reaction = a reaction involving the transfer of a proton



Bronsted – Lowry theory of acids and bases

Acid Base



Base Acid



Bronsted – Lowry theory of acids and bases

Acid Base



Base Acid

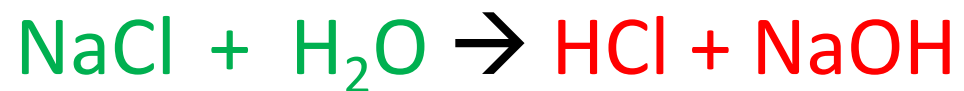


Bronsted – Lowry theory of acids and bases

Acid Base



Base Acid



Acid Base Base Acid



Bronsted – Lowry theory of acids and bases

Acid Base



Base Acid



Acid Base Base Acid



Bronsted – Lowry theory of acids and bases

Acid Base



Base Acid



Acid Base Base Acid



Bronsted – Lowry theory of acids and bases

Acid Base



Base Acid



Acid 1 Base Base 1 Acid



Bronsted – Lowry theory of acids and bases

Acid Base



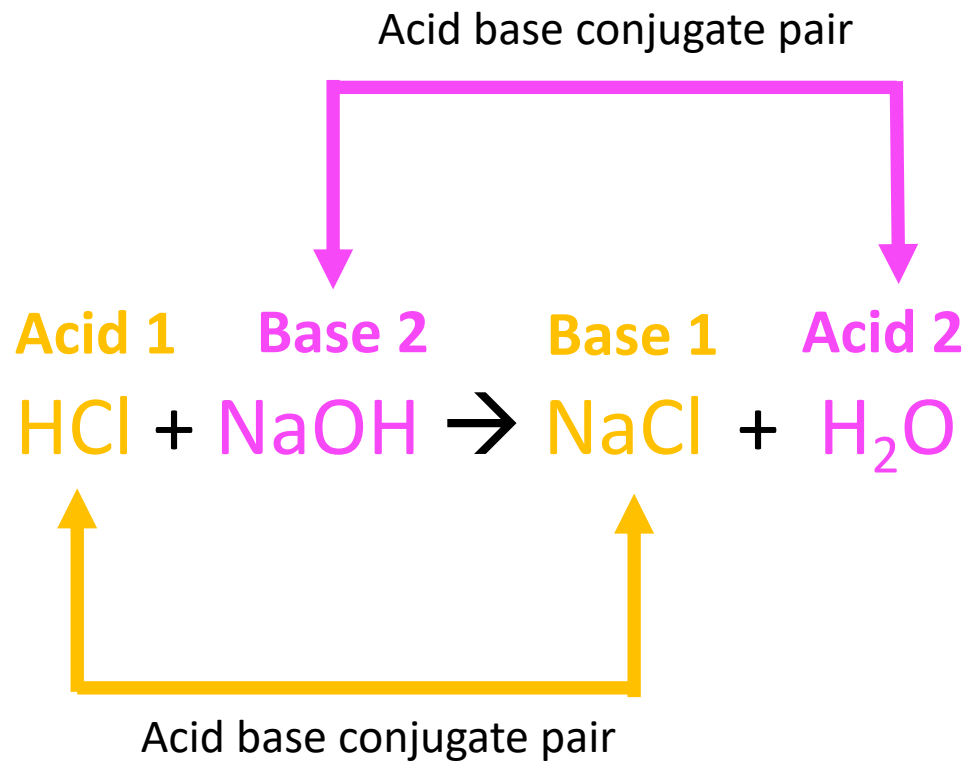
Base Acid



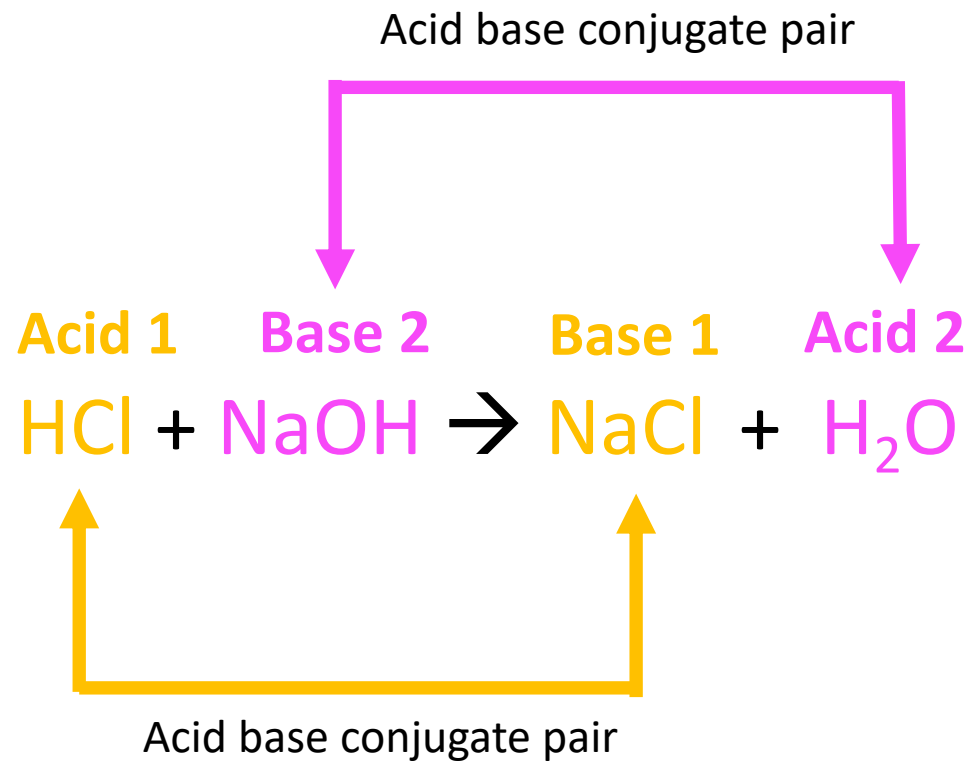
Acid 1 Base 2 Base 1 Acid 2



Bronsted – Lowry theory of acids and bases



Bronsted – Lowry theory of acids and bases



Conjugate acid base pairs are chemicals which have the same formula except that the acid has an extra H⁺ in its formula.

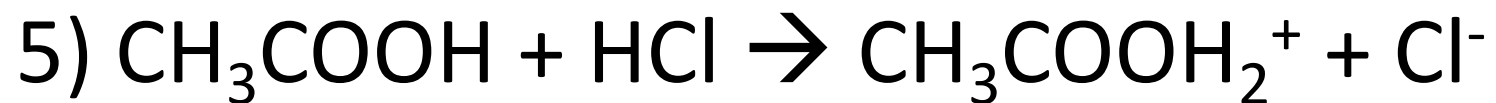
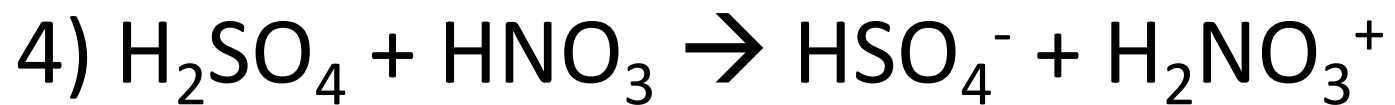
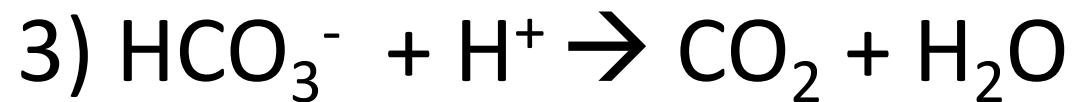
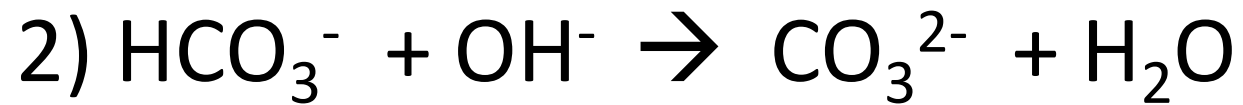
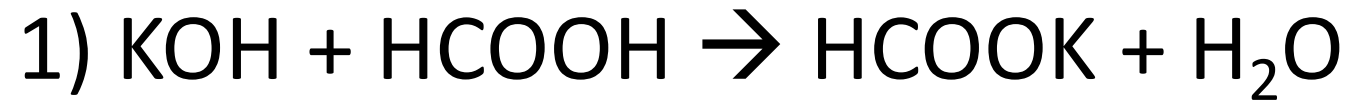
Bronsted – Lowry theory of acids and bases

Conjugate acid	Conjugate base
HCl	Cl ⁻
H ₂ O	OH ⁻
HNO ₂ ²⁺	NO ₂ ⁺
H ₃ O ⁺	H ₂ O



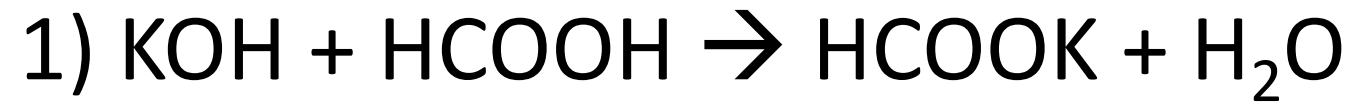
Check up quiz

Identify the acid-base pairs for the following reactions

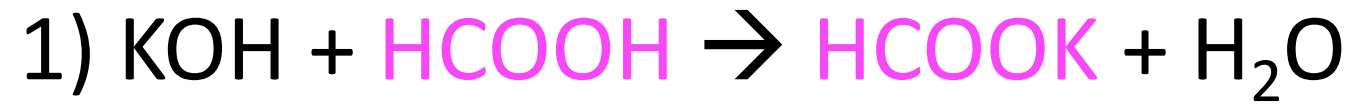


Answers
coming up
now...

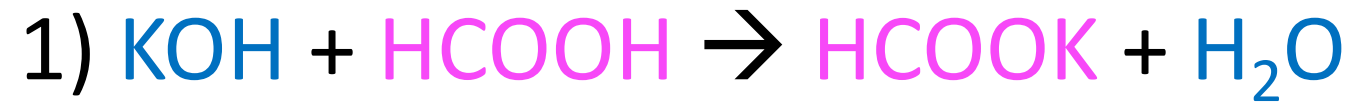
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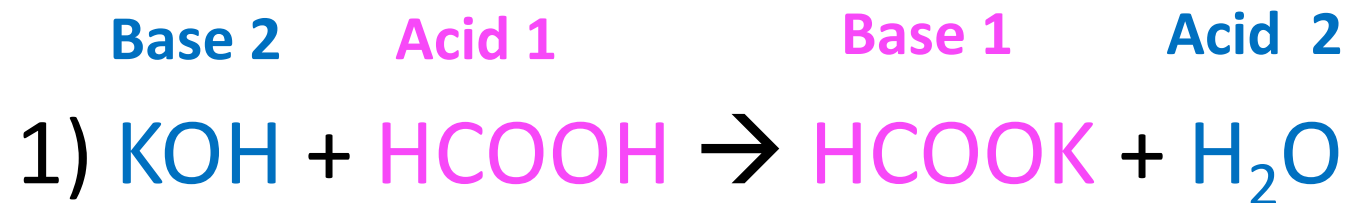
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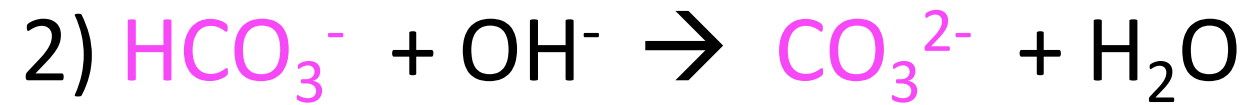
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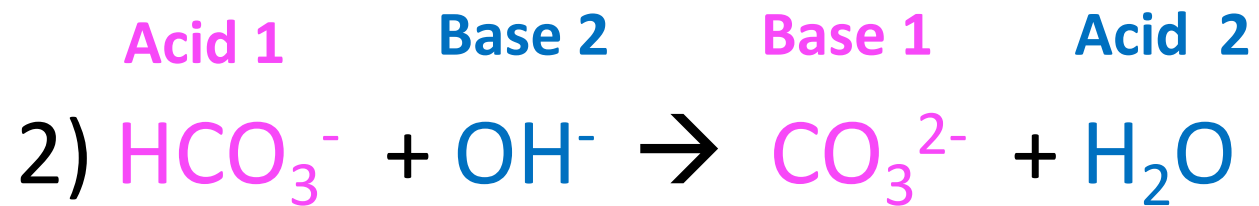
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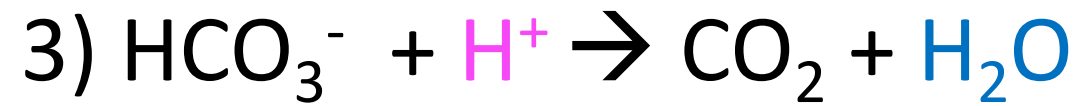
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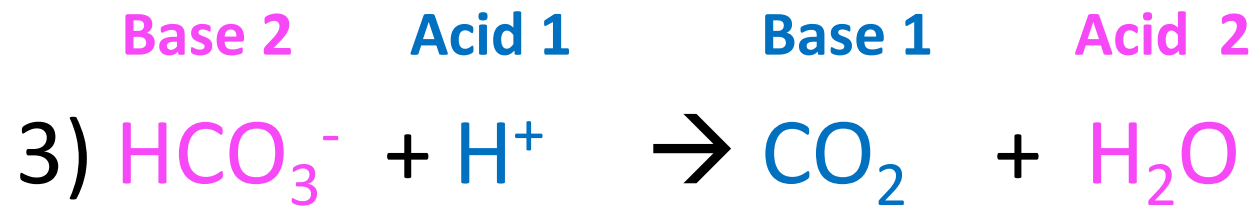
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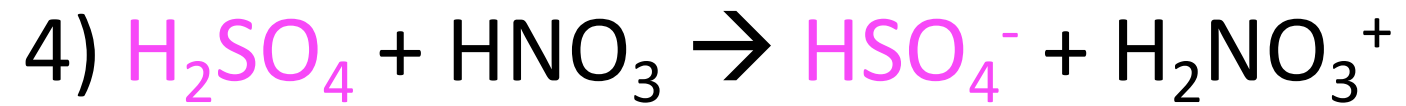
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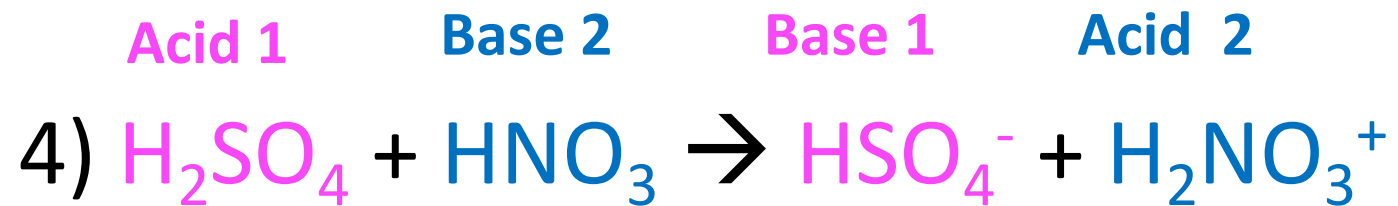
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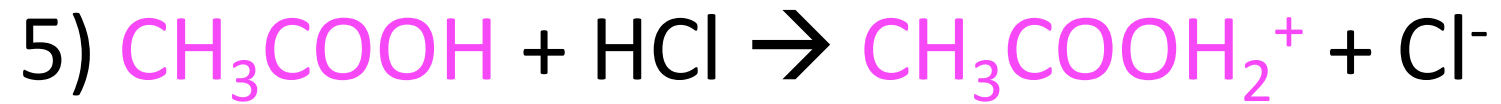
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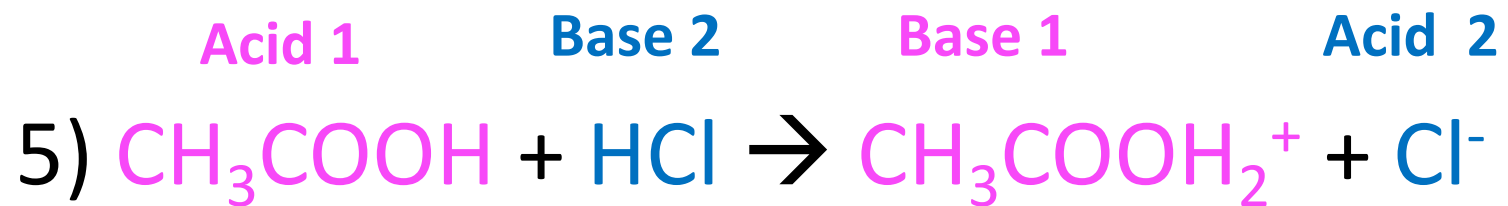
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